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K_b K_w pH pOH pKa

pKb H^+ OH^-

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pH, pOH, H_3O^+ , OH^- ,

K_w , K_a , K_b , pKa, and

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Arrhenius Model 1.

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Acids produce

hydrogen ions in aqueous solutions 2.

Bases produce hydroxide ions in aqueous solutions B.

Bronsted-Lowry Model 1. Acids are proton donors 2.

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Acidic*acid* ** low*

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Properties Of Acids

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Introduction.

Originally the terms acid and base referred to taste. The practice of classifying substances according to their acidic (sour) or basic (alkaline or bitter) properties dates back to ancient

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times. An acid was something with a sour taste, such as lemon juice, and a base was something with a bitter taste, such as tonic water.

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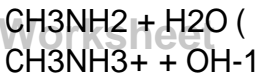
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1) Using your knowledge of the Brønsted-Lowry theory of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair: a) $\text{HNO}_3 + \text{OH}^-$ ($\text{H}_2\text{O} + \text{NO}_3^-$). HNO_3 and NO_3^- make one pair OH^- and H_2O

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make the other. b)



Acid and Base

Worksheet - Answers

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Bases an acid that is

composed of H, O,

and a third element.

The H & O are

covalently bonded

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and are called the hydroxyl group.

Property of bases.

taste bitter, feel

slippery, change

color of indicators,

react with acids to

produce a salt and

water neutralization

rxn, are electrolytes.

Arrhenius or

traditional base.

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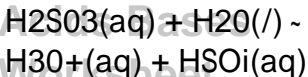
SHORT ANSWER

Answer the following questions in the space provided. 1. a.

Write the two equations that show the two-stage ionization of sulfurous acid in water. stage 1:

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b. Which stage of ionization usually produces more ions? Explain your answer.

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